

Multi-step microfluidic reactor for the synthesis of hybrid nanoparticles

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Received 10 March 2020, revised 21 April 2020

Accepted for publication 28 April 2020

Published 8 June 2020

Abstract

We introduce a multi-step microfluidic reactor for the synthesis of hybrid nanoparticles. As part of this study, nanoparticles composed of chitosan and iron-oxide are synthesized at room temperature by forming sequential droplets of reagents in the microreactor followed by merging and mixing them in a step by step process. The obtained nanoparticles were characterized by transmission electron microscopy, Fourier transform infrared spectrometry, and energy dispersive x-ray analysis. Results were compared with nanoparticles of the same composition synthesized with batch-wise conventional techniques. As a result, the obtained nanoparticles showed better size distribution. This microfluidic device can be used for the synthesis of other types of nanoparticles that require multi-step procedures.

Keywords: microfluidics, microreactor, droplet-based flow, droplet merging, nanoparticle synthesis, hybrid nanoparticles

(Some figures may appear in colour only in the online journal)

1. Introduction

Hybrid nanomaterials are composed of two or more types of materials that show multifunctional behavior which makes them promising for applications in branches of science, engineering and medicine [1–3]. It is very crucial to synthesize these materials monodispersely as their properties are highly dependent on their size, shape and composition. Mostly nanomaterials are synthesized with conventional batch-wise methods, however these techniques have limitations in maintaining uniform reaction conditions such as temperature, concentration, residence time and therefore they usually result in polydisperse size distributions. On the other hand microfluidic reactors are shown to provide precise control over these parameters resulting in much monodisperse products [4–6] with reduced sample usage and minimized exposure to toxic materials [7].

Synthesis of nanoparticles using microfluidic systems has been studied earlier and there are also studies that focused on the synthesis of hybrid nanoparticles [8]. Hybrid materials usually require multi-step synthesis processes and microfluidic reactors have the ability to perform those steps separately in a controlled environment and thereby can produce

homogeneous size distribution [7]. Among the studies about microfluidic synthesis of composite and hybrid nanoparticles, a number of them focused on obtaining particles with a polymer component. For instance, Zhang *et al* synthesized PDMS coated gold nanoparticles [9]; Hwang *et al* synthesized magnetic hydrogel microparticles [10]; Abou-Hassan *et al* utilized a multi step microreactor and coated iron oxide with silica with an average size of 50 nm [11]; Valencia *et al* obtained lipid-quantum dot hybrid nanoparticles in a continuous flow microreactor [12]; Kumar *et al* produced dextran coated iron oxide nanoparticles for MRI applications [13]; Yang *et al* used a semi-microfluidic platform to synthesize chitosan coated silver nanoparticles where droplets of silver nitrate and chitosan were generated within the microfluidic device whereas the reaction was completed outside of the reactor in a sodium hydroxide bath [14]; Köhler *et al* synthesized SERS active sensor polyacrylamide microparticles with silver nanoparticles [15]; Feng *et al* obtained lipid-polymer nanohybrids [16]; Hassan *et al* formed Janus particles composed of silica and gold within a droplet-based microreactor [17]; and Li *et al* used a flow focusing device to produce lipid-polymer hybrid nanoparticles [18]. There are also studies that focused on the synthesis of metal–metal and metal–metal

oxide nanohybrid particles. Co@Au core shell nanohybrids [19, 20], Fe@Fe₃O₄ nanoparticles [21], PtSn alloyed nanoparticles [22], Pt-Pb core shell nanohybrids [23], Ag-ZnO nanohybrids [24], and Ag-CoFe hybrid nanoparticles [25] are among these synthesized particles.

This paper focuses on synthesizing hybrid nanoparticles composed of chitosan and iron oxide by a multi-step process carried in a novel microfluidic reactor. Chitosan, a biopolymer, is a well known material for its biocompatibility, low toxicity, biodegradability and stability which makes it very suitable for medical and pharmaceutical applications, especially in drug delivery [26–29]. When it is synthesized in nanometer dimensions, due to its small size and pH responsive chemistry it can penetrate into cells to release the drug it carries [28]. On the other hand, magnetic nanoparticles behave as a single domain material and therefore when there is an applied external magnetic field, they act like a paramagnet. Therefore, nanoparticles composed of both a magnetic material and chitosan are promising for applications such as drug delivery and magnetic resonance imaging making them both a therapeutic and a diagnostic tool [30].

Chitosan coated iron-oxide nanoparticles have been synthesized previously with batch-wise synthesis methods. Unsoy *et al* synthesized chitosan coated iron oxide nanoparticles and investigated the effect of the concentration of ammonia solution on the size of the particles [31]. The diameter of the synthesized nanoparticles were ranging between 2 nm and 8 nm. Safari *et al* obtained chitosan coated iron-oxide nanoparticles using a similar approach with an average diameter of 10 nm and used these particles as catalyst for the synthesis of phenytoin derivatives [32]. Zarnegar *et al* synthesized and reported the use of composite Fe₃O₄ @ chitosan nanoparticles as catalyst for the synthesis of 2, 4, 5-trisubstituted imidazoles [33]. He *et al* described a single step and two step process for the synthesis of chitosan decorated iron-oxide nanoparticles with an average size of 300 nm and reported their performance for the removal of furfural from aqueous solutions [34].

There is still a need to improve the size distribution of hybrid nanoparticles composed of chitosan and iron-oxide in literature. In order to obtain monodisperse particles with desired sizes a new microfluidic platform is introduced in this article that is capable of generating alternating droplets of reagents followed by merging and mixing them in controlled amounts. This multi-step microreactor can also be used in the synthesis of other hybrid nanomaterials that require processing of more than two reagents.

2. Experimental

2.1. Materials

Chitosan, ferric chloride hexahydrate (FeCl₃ · 6 H₂O) and ferrous chloride tetrahydrate (FeCl₂ · 4 H₂O) were obtained from Sigma Aldrich. Synthesis was performed by following a similar method described by He *et al* [34]. Solutions were prepared by using deionized (DI) water. 0.75 g of chitosan was mixed and stirred in 100 ml of 0.3% acetic acid aqueous solution for 30 minutes (dispersed phase 1). 1.09 g of FeCl₃ · 6 H₂O

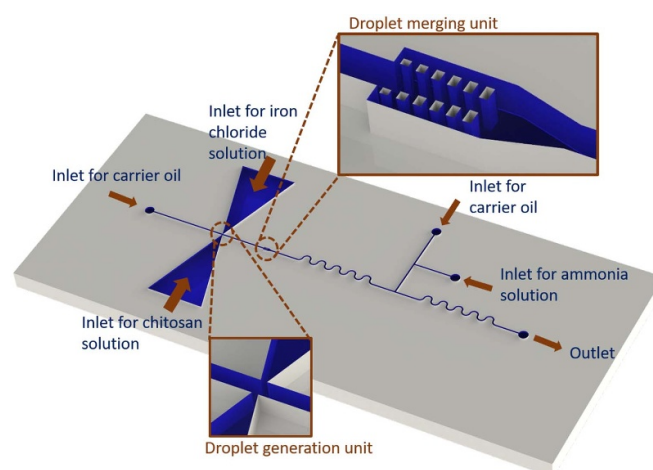


Figure 1. Schematic of the microreactor.

and 0.4 g of FeCl₂ · 4 H₂O were mixed and stirred in 100 ml of DI water for 30 minutes (dispersed phase 2). As a third reagent solution, 10% aqueous ammonia solution is used (dispersed phase 3). Silicon oil was used as the continuous phase in the channels.

2.2. Fabrication of the microfluidic device

Microfluidic device was fabricated out of polydimethylsiloxane (PDMS) by using conventional soft-lithography technique. The master mold was made out of negative photoresist (SU-8 2050, Microresist Technology) on a polished silicon wafer. Microfluidic channels obtained in PDMS were bonded to a glass slide after oxygen plasma treatment.

2.3. Experimental set-up

Reagents were delivered to the microreactor by using syringe pumps. An inverted microscope with a camera was used to observe and record the flow.

3. Design of the microfluidic device

The microfluidic device is designed to carry out a multi-step synthesis by utilizing a droplet-based flow. First, alternating droplets of dispersed phase 1 and 2 are generated at the double T-junction. Then these two reagents are merged and mixed. At the final step, a droplet of dispersed phase 3 is added to the mixture to precipitate iron-oxide and complete the reaction. These steps are explained in detail below.

3.1. Alternating droplet generation

In the synthesis of hybrid nanoparticles of iron oxide and chitosan, three reagent solutions need to be mixed. In order to control the mixing ratio precisely and to prevent channel contamination, a droplet-based flow is utilized in channels. First, droplets of iron chloride and chitosan solutions are generated alternatively at a T-junction with tapered inlets located across

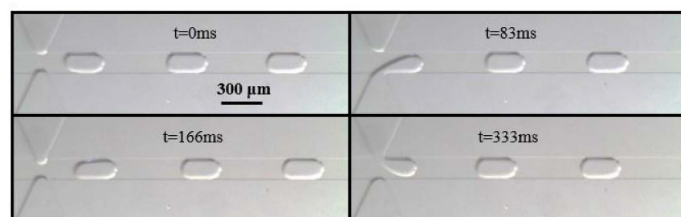


Figure 2. Alternating reagent droplet generation using tapered inlets. Tapered geometry controls the pressure at the junction which enables the formation of droplets in a sequence.

Table 1. Flow rate condition for optimum results.

Reagent	Flow rate ($\mu\text{l min}^{-1}$)
Carrier oil	1.5
Iron chloride solution	0.17
Chitosan solution	0.12

from each other. Saqib *et al* showed the taper angle assures an alternating droplet generation from opposing inlets due to the control of Laplace pressure and optimized the angle of taper as 25° to obtain stable pattern generation and uniform droplet size and spacing [35]. As chitosan and iron chloride solution need to be mixed in 1:1 ratio, it was made sure that the pattern of alternating droplets is repetitive. As shown in figure 2, first the lower stream (dispersed phase 2) enters the main channel, forms a droplet and afterwards the upper stream (dispersed phase 1) generates the next droplet. This process continues until flow is stopped.

In adjusting the flow rates of reagents, viscosity has to be taken in account as well. As chitosan solution has higher viscosity than iron chloride solution, it is essential to keep flow rate of iron chloride solution to be high in order to achieve droplet generation with the same volume. Table 1 summarizes the flow rates for which the best results were obtained.

These flow rates not only define the pattern but are also essential for homogeneous droplet size formation, thereby influencing the success of the reaction. It is to be noted that for alternating droplet generation, the flow ratio plays a vital role rather than the individual flow rates. But as the alternating droplet generation platform is linked with droplet merging platform, which is influenced by carrier oil flow rate, optimum values were finalized after observing the effect on both the droplet generation and merging.

3.2. Droplet merging

After generating the alternating droplets of two reagents, two consecutive droplets of iron chloride and chitosan solutions are merged to initiate the synthesis. An expanded channel with pillars ensures the merge of the same number of droplets in a repetitive and controlled way without disturbing the rest of the flow in the system [36]. This passive geometry prevents a third droplet to join the other two droplets which assures the same concentration in each merged droplet.

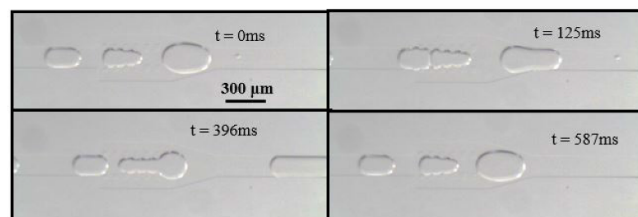


Figure 3. Passive merging of alternating droplets in pillar induced structure.

In the merging unit, the width of the pillared channel increases while the width of the pillars increases as well. Increase in the cross sectional area causes a decrease in the flow rate which also slows down the motion of droplets to an extent where they make contact and eventually merge.

Pillar induced merging geometry is optimized for merging to take place with only two consecutive droplets without any side flow. If a part of droplet flows from sides it means that the concentration of reagent is changed and the reaction parameters are no longer the same. Hence it was essential to optimize the dimensions. For this purpose the width of the channel was designed such that it increased from $150 \mu\text{m}$ to $335 \mu\text{m}$ and the width of the pillars were increased by $10 \mu\text{m}$ for first four pillars and then by $5 \mu\text{m}$ for the rest. In this case two consecutive droplets merged successfully without any side flows.

Figure 3 shows the merging of two consecutive reagent droplets. Merged droplet later passes through serpentine channels where mixing takes place due to the circulating flow profile within droplets [37]. The height of microchannel throughout the microfluidic device is $100 \mu\text{m}$. In the serpentine channel the width is kept uniform at $150 \mu\text{m}$ which is the same throughout the device except the merging unit. The outer radius is $450 \mu\text{m}$ and the inner radius is $300 \mu\text{m}$, resulting with a mean radius of $375 \mu\text{m}$. There are 11 arcs after the merging geometry with a length of approximately 1.3 cm and 12 arcs after the addition of ammonia with length of approximately 1.4 cm

3.3. Addition of the third reagent

In the final step of the synthesis, ammonia solution is added as a third reagent to precipitate iron oxide. First, a droplet of ammonia solution is generated at a T-junction and later it is fused with the droplet composed of chitosan and iron chloride

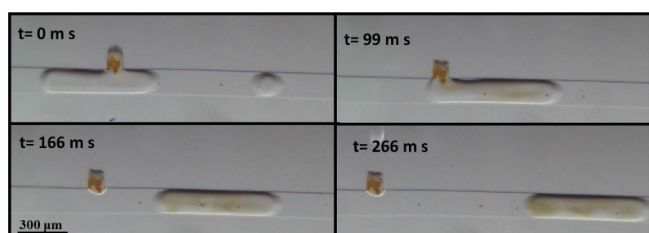


Figure 4. Merging of ammonia droplet to initiate reaction.

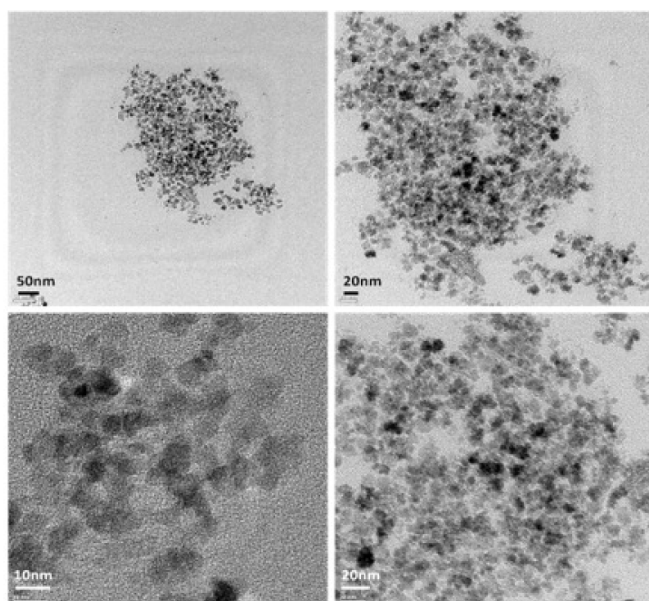


Figure 5. TEM images of synthesized nanoparticles.

solution. As soon as the ammonia solution is added, oxidation of iron is observed with the color change within droplets as shown in figure 4.

The amount of added ammonia solution is essential in the size of the synthesized nanoparticles. Unsoy *et al* showed that increasing concentration of ammonia solution decreases the size of nanoparticles [31].

4. Results and discussion

4.1. Results

To understand their both physical characteristics and chemical composition, nanoparticles synthesized in the microreactor were characterized by various analytical tools including transmission electron microscopy (TEM), energy dispersive x-ray analysis (EDX) and Fourier transform infrared spectroscopy (FTIR).

Images of the particles were recorded with TEM and average size of nanoparticles were measured as 8.5 nm with a standard deviation of 1.3 nm after image processing of the data. TEM images are shown in figure 5. The size of these particles are smaller than the ones synthesized by batch techniques reported in literature [31, 34] mainly due to having a

Table 2. EDX data of synthesized nanoparticles.

Element	Percentage
Fe	44.19
O	55.8

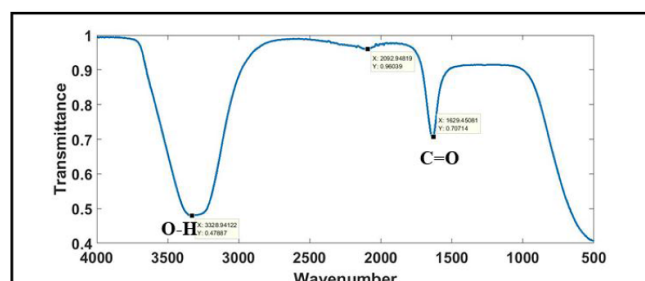


Figure 6. FTIR spectra of synthesized nanoparticles.

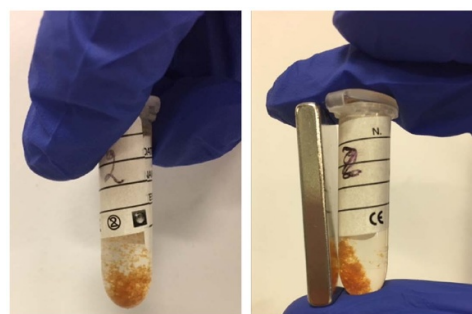


Figure 7. Nanoparticles without and with the presence of a magnet.

synthesis constrained in droplets as small as 19 nL. On the other hand, an improved size distribution is one of the major advantages of using microfluidic systems which is a result of maintaining a uniform concentration and residence time for all reagents. These images also show that particles are spherical; on the other hand, since ligands were not used during the synthesis there is agglomeration.

EDX analysis was performed to confirm the existence of iron-oxide component in the nanoparticles. Only the elements Fe and O were selected to be scanned in the system. Since the grid structure used for holding the sample was made out of carbon, scanning of carbon was excluded from the analysis as it would not result in an accurate conclusion for chitosan. EDX data is presented in table 2 and shows the percentage distribution of Fe to O. In order to confirm the presence of chitosan, FTIR analysis was made and the result is shown in figure 6. Successful synthesis of hybrid chitosan and iron-oxide nanoparticles is confirmed by the appearance of characteristic bands at 3328.94 cm^{-1} indicating the bonding of N-H group with O-H group in chitosan and 1629.45 cm^{-1} indicating the C = O of alcoholic group in chitosan.

The synthesized nanoparticles were also magnetic as shown in figure 7.

4.2. Discussion

The analytical data obtained from TEM, FTIR and EDX confirm the successful synthesis of nanoparticles composed of chitosan and iron oxide. Compared to the batch synthesized chitosan coated iron-oxide nanoparticles reported by He *et al* [34], the nanoparticles produced in the microreactor were two orders of magnitude smaller in size. A size distribution data was not reported in [34], therefore a direct comparison is not possible to be made; on the other hand, similar particle agglomeration was observed as the synthesis was carried out without any surfactant or ligand in both cases. Compared to another batch method reported for the synthesis of chitosan coated iron-oxide nanoparticles by Unsoy *et al* [31], it is observed that size distribution was improved with the microreactor. Unsoy *et al* also observed agglomeration of nanoparticles.

5. Conclusion

Synthesis of hybrid chitosan with iron-oxide nanoparticles has been realized by using a microfluidic platform which resulted in a narrower size distribution and reduced residence time compared to the batch-wise methods. Alternating droplet formation, controlled merging and step by step mixing was carried out in a single microfluidic chip. Dimensions of the device were optimized and flow rate of reagents were adjusted to meet the requirements of the synthesis recipe. The designed microfluidic device can be used to produce other materials that require a multi-step synthesis protocol.

Conflicts of interest

Authors declare no conflict of interest.

Acknowledgments

This work was funded by the Science Academy Young Scientists Award (BAGEP 2016). We also thank Eliza Sopubekova for her assistance in FTIR and Mr Mustafa Guler for his assistance in TEM analysis.

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